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book, where the pressure increases from 0.023 bar to 300 bars, the enthalpy change from 1 to 2 is $\Delta H_{12} \approx 300 \cdot 10^5 \cdot 10^{-3} = 3 \cdot 10^4 \text{ J kg}^{-1} = 30 \text{ kJ kg}^{-1}$ (19)

Schroeder's example uses $P_1 = 0.023 \text{ bar}$ (giving $T_1 = 20 \text{ C}$), $P_2 = 300 \text{ bars}$ and $T_3 = 600 \text{ C}$. This changes the efficiency given by Schroeder to $e=1$

STEAM ENGINES; THE RANKINE CYCLE An Introduction to ... ISLAMABAD INSTITUTE ISLAMABAD INSTITUTE Thermodynamic identity (Schroeder 3.4) $dU = TdS - pdV$ (1.5) Total differentials When we have a total differential, the things in front of the dS, dV , etc are partial derivatives. $T = \frac{\partial U}{\partial S} \bigg|_V - p = \frac{\partial U}{\partial V} \bigg|_S$ (1.6) Also, you can integrate along a path using a total differential, and can do linear algebra using total differential equations, e.g. substituting one for another.

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