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mass. Recall from Chapter 3 that the law states that matter is neither created nor destroyed in a chemical reaction. Chapter 11: Stoichiometry Tomorrow's answer's today! Find correct step-by-step solutions for ALL your homework for FREE! Chemistry Textbooks :: Homework Help and Answers :: Slader 1 point for answer (ii) What is the limiting reactant for the reaction? Justify your answer with calculations.

$$n_{\text{Fe}_2\text{O}_3} = \frac{15.39 \text{ g Fe}_2\text{O}_3}{159.7 \text{ g Fe}_2\text{O}_3/\text{mol Fe}_2\text{O}_3} = 0.0964 \text{ mol Fe}_2\text{O}_3$$
$$n_{\text{CO}} \text{ required} = 0.0964 \text{ mol Fe}_2\text{O}_3 \times 3 \text{ mol CO} / 1 \text{ mol Fe}_2\text{O}_3 = 0.289 \text{ mol CO}$$

required to completely react with 0.0964 mol Fe₂O₃ AP*

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Section Quiz: Introduction to Stoichiometry In the
space provided, write the letter of the term or phrase

that best completes each sentence or best answers each question. 1. Stoichiometry is the branch of chemistry that deals with elements in

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